

Comparative Study of Three Cathodic Electron Acceptors on the Performance of Mediatorless Microbial Fuel Cell

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Abstract: Microbial Fuel Cells (MFC) are capable of converting the chemical energy stored in the chemical compounds in a biomass to electrical energy with the aid of microorganisms. A better understanding of the relationship between the microbiological and electrical aspects of fuels cells is needed prior to successful MFC application. MFC constructed employing low-cost materials (salt bridge, graphite electrodes) without addition of any toxic mediators and using glucose (3 g L^{-1}) as substrate and three different electron acceptors potassium permanganate (KMnO_4), potassium ferricyanide ($\text{K}_3[\text{Fe}(\text{CN})_6]$) and potassium dichromate ($\text{K}_2\text{Cr}_2\text{O}_7$) in the cathode chamber without pH adjustment was assessed in the present study. The results showed that MFC performance with KMnO_4 was superior than $\text{K}_3[\text{Fe}(\text{CN})_6]$ and $\text{K}_2\text{Cr}_2\text{O}_7$. Maximum voltage generation (OCV) and power density for KMnO_4 were 1.04 V and 7.29 mW m^{-2} . Performance of MFC with $\text{K}_3[\text{Fe}(\text{CN})_6]$ and $\text{K}_2\text{Cr}_2\text{O}_7$ was much lower with a OCV and power density of 0.71 V; 0.92 mW m^{-2} and 0.56 V; 0.79 mW m^{-2} , respectively. Demonstrating that KMnO_4 can be used as an effective cathodic electron acceptor to enhance power output on MFC.

Key words: Cathodic electron acceptor, electrical energy, microbial fuel cell, potassium permanganate, power density, voltage

INTRODUCTION

Recently, Microbial Fuel Cells (MFCs) have emerged as a promising yet challenging technology. MFCs are considered to be an alternative sustainable technology to meet increasing energy needs, simultaneously accomplishing efficient wastewater treatment and electricity generation (Aelterman *et al.*, 2006; Lu *et al.*, 2009; Ahn and Logan, 2010). In this technology, bacteria act as biocatalysts to oxidize organic and inorganic matter and generate current (Oh *et al.*, 2004; Li *et al.*, 2008). Electrons produced by bacteria from these substrates migrate to the anode (negative terminal) in the anodic chamber and flow to the cathode (positive terminal) through a conductive material and a resistor, or operated under a load (i.e., producing electricity that runs a device). Electrons can be transferred to the anode by electron mediators or shuttles (Rabaey *et al.*, 2005) by direct

membrane associated electron transfer (Bond and Lovley, 2003) or by so-called nanowires (Reguera *et al.*, 2006) produced by the bacteria or perhaps by other as yet undiscovered means.

For a long period of time, MFCs remained a scientific curiosity because of their limited efficiency (Schröder, 2007). At present, power outputs in MFC are being increased due to the integration of both microbial and electrochemical aspects, important hurdles have been tackled (Oh *et al.*, 2004; Rabaey *et al.*, 2005). Nevertheless, several bottlenecks still exist, each requiring an appropriate strategy development.

Previous reports have shown that the performance of a MFC is determined by several factors (Pham *et al.*, 2004). They are: microbial activities oxidizing fuels in the anode, electron transfer from microbial cells to the anode, proton transfer from the anode to cathode through the selective membrane, resistance of the circuit used to

transmit electrons from the anode to the cathode and cathode reaction. Among them, oxygen consumption is considered to be the most serious factor to limit the performance of MFC. Increasing efforts have been devoted to optimizing the reduction of the electron acceptor on the cathode electrode surface. So far, oxygen is still widely used as the electron acceptor in most MFC applications due to its availability in the environment and its high redox potential (Cheng *et al.*, 2006; Clauwaert *et al.*, 2007; Li *et al.*, 2009). However, the main disadvantages is the poor contact between gaseous oxygen and cathode electrode as well as the slow rate of oxygen reduction on the surface of carbon electrodes hinder MFC application (Li *et al.*, 2009). In order to eliminate oxygen limitation, the dissolved oxygen concentration in the cathode should be kept high, which requires increased power consumption and results in more oxygen diffusion into the anode. Diverse cathode electron acceptors with improved catalytic activity have been proposed to overcome this problem (Pham *et al.*, 2004; Oh *et al.*, 2004; Tartakovsky and Guiot, 2006; You *et al.*, 2006).

In the study, a two-chamber MFC was constructed, using low cost construction material, reused graphite rod electrode, two plastic bottle linked with salt bridge. Three cathodic electron acceptors were assessed (potassium dichromate, potassium permanganate and potassium ferricyanide), comparison between the three electron acceptors was established by means of voltage generation (OCV open circuit voltage) and power density.

MATERIALS AND METHODS

Experimental set-up: The two-chamber MFC were constructed with two plastic bottles with a volume of 1000 mL each. Plastic bottles were perforated and attached an inner screw thread joint to each bottle. An external thread joint (4 in length) was used to construct the agar salt bridge (agar 3%, KCl 100 Mm). Graphite rods (\varnothing 0.8×4 cm length, surface area 10.55 cm²) were used as electrodes for both chambers and connected using coated copper wire 22 gauge and sealed with silicon. Anodic chamber was inoculated with 250 mL of anaerobic sludge obtained from a UASB reactor and 500 mL of medium containing glucose as sole carbon source (3 g L⁻¹ = 850 mg COD L⁻¹) in a mineral medium Na₂H₂PO₄·2H₂O (0.53 g L⁻¹), NH₄Cl (0.3 g L⁻¹), MgCl₂·2H₂O (0.1 g L⁻¹), CoCl₂·6H₂O (0.08 g L⁻¹), NiCl₂·6H₂O (0.02 g L⁻¹) and ZnCl₂ (0.001 g L⁻¹). Cathodic solutions (20 Mm each) were prepared with analytical-grade potassium permanganate (KMnO₄), potassium ferricyanide (K₃[Fe(CN)₆]) and potassium dichromate (K₂Cr₂O₇) chemicals. No mediators

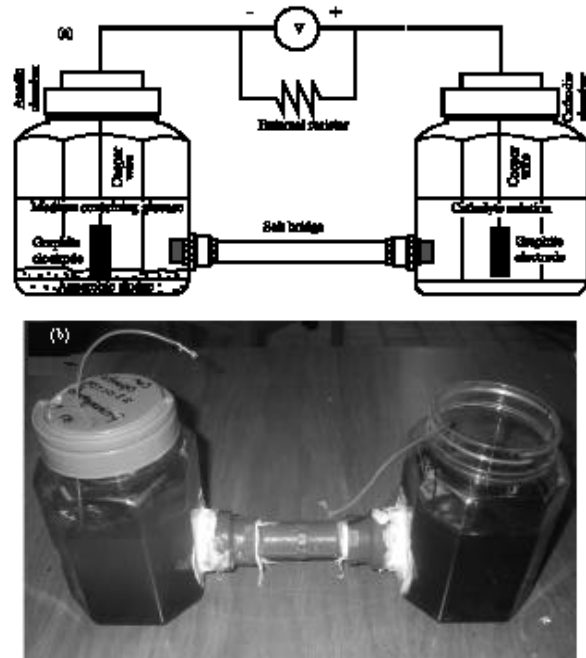


Fig. 1: Schematic (a) and picture (b) of the two-chamber MFC used on this study

were used in the anodic chamber of the MFC. Evaluation of MFC was conducted at room temperature (18-23°C). The schematic and picture of two-chamber MFC are illustrated in Fig. 1a and b.

Measurements and calculations: Open Circuit Voltage (OCV) was measured every 60 min with a digital multimeter (FLUKE 287), external resistor (4.7×5 K Ω) was connected to determine current as describe by Schaeztle *et al.* (2009). The external resistor variation was in a range of 10 Ω -6.2 M Ω , to determine the maximum power density of the two-chamber MFC (mW m⁻²) according to Eq. 1 (Logan *et al.*, 2006), where V (volts) denotes the voltage, I (mA) the current and A (m²) is the surface area of the electrode. The glucose overall removal as COD was determined as describe by Rodriguez-Martinez *et al.* (2005):

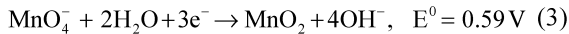
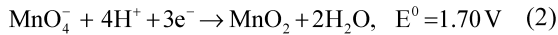
$$P = \frac{VI}{A} \quad (1)$$

RESULTS AND DISCUSSION

Electricity generation in a two-chamber MFC using potassium permanganate, potassium ferricyanide and potassium dichromate as electron acceptors: the two-chamber MFC were operated in batch mode by adding glucose as substrate (3 mg L⁻¹ glucose \approx 854 \pm 2 g O₂ L⁻¹)

in the anode chamber and three different electrolyte solution were compared as electron acceptors in the cathode chamber, potassium permanganate (KMnO₄), potassium ferricyanide (K₃[Fe(CN)₆]) and potassium dichromate (K₂Cr₂O₇).

Figure 2 shows the voltage (OCV) generated for each case after 814 h. During the first 48 h OCV increased rapidly for all cases, the highest maximum OCV (1.07 V) was achieved in MFC using KMnO₄ as electron acceptor. While, K₃[Fe(CN)₆] and K₂Cr₂O₇ only generated 0.71 and 0.56 V, respectively (Table 1). No pH decreasing adjustment were made as it has been described in early reports (Zhao *et al.*, 2005; You *et al.*, 2006; Li *et al.*, 2008, 2009) that low pH values lead to an increase of cathode potential, since low pH induces a higher radical yield. As described by You *et al.* (2006) Eq. 2 and 3 illustrate the fact that in acidic condition, MnO₄⁻ has a much higher oxidation potential than in alkaline conditions.



Power and current density: By varying the circuit resistance from 10Ω-6.2 MΩ, a maximum current density of 6.6 μA cm⁻² for the case of KMnO₄ as electron acceptor was obtained, being more than an order of magnitude higher than the other cases studied (K₃[Fe(CN)₆] and K₂Cr₂O₇) with a maximum power density of 7.29 mW m⁻² (Table 1 and Fig. 3). Similar results were reported by Mohan *et al.* (2008) were maximum power density achieved was 9.3 mW m⁻² using methylene blue (0.03 mM as optimum concentration) as mediator, with the difference that no mediators were used on the present study as mentioned above. Mediators trap and draw the electrons from membrane-bound reactions involving proton and electron transfer and supply those electrons to the anode electrode which is at a lower potential than the mediator itself. Since the surface area of the electrode is not sufficient to cover all of the growing bacteria, mediator acts as electron shuttles. However, the toxicity and instability of synthetic mediators limit their applications in MFCs.

The maximum power density and current reported here are modest compared with MFC using a proton exchange membranes, due to higher internal resistance of the salt bridge system as describe by Min *et al.* (2005), where comparison between MFC with two different selective membranes were conducted (proton exchange

Table 1: Electricity generation (OCV), power outputs and COD removal for KMnO₄, K₃[Fe(CN)₆] and K₂Cr₂O₇ as electrolyte cathodic solution in a two-chamber MFC after 814 h

Cathodic acceptor	OCV (volts)	Power density (mW m ⁻²)	COD removal (%)
KMnO ₄	1.04	7.29	95±1.5
K ₃ [Fe(CN) ₆]	0.71	0.92	95±1.5
K ₂ Cr ₂ O ₇	0.56	0.79	95±1.5

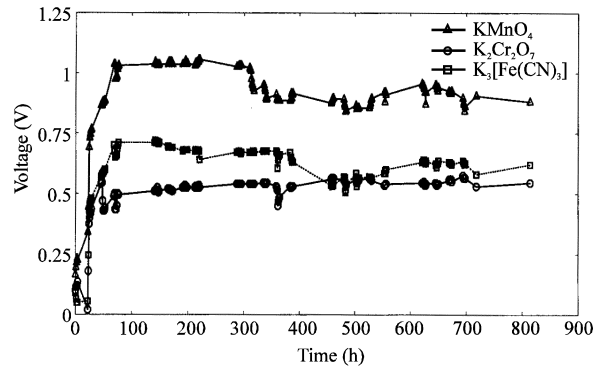


Fig. 2: Comparison of voltage (OCV) generation in two-chamber MFC using three different cathodic electron acceptors (KMnO₄, K₃[Fe(CN)₆], K₂Cr₂O₇, 20 mM each); (3 mg L⁻¹ glucose ≈ 854±2 g O₂ L⁻¹)

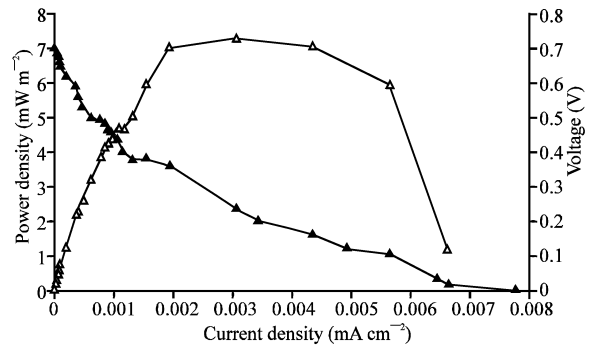


Fig. 3: Polarization and power curves in two-chamber MFC using KMnO₄ (20 mM), as cathodic electron acceptor (3 mg L⁻¹ glucose ≈ 854±2 g O₂ L⁻¹). K₃[Fe(CN)₆] and K₂Cr₂O₇ polarization and power curves are not shown on Fig. 3 due to the great difference compared to KMnO₄

membrane vs. salt bridge), the low power output by the MFC with a salt bridge is directly attributed to the high internal resistance (19920±50Ω) compared to the low internal resistance by the MFC with proton exchange membrane (1286±1Ω). However, the results obtained in this study are found to be very promising for a salt bridge two-chamber MFC.

KMnO₄ can be used as an effective cathodic electron acceptor in MFC, achieving a higher power density compared with other electron acceptors such as K₃[Fe(CN)₆] and K₂Cr₂O₇. Additionally KMnO₄ is environment-friendly oxidant in comparison.

CONCLUSION

The influences of cathodic electron acceptor solutions (cathode solutions) on the power generation of two-chamber MFC were investigated in this study. The results show that optimizing power generation in MFC requires maximizing proton transport rates, the OCP value depends on the type of electron acceptor used.

As KMnO₄ demonstrated to be a good option as cathodic electron acceptor, producing higher OCV and power density than K₃[Fe(CN)₆] and K₂Cr₂O₇. In order to improve the power generation of MFCs, it is critical to elucidate the limiting factor for power generation.

One efficient approach to understand the limiting factor is to systematically compare these factors for MFC performance, factors such as distance between the electrodes, surface area of the electrodes and different selective membrane to achieve a substantial reduction on internal resistance of the system will be studied further to check the efficiency of the same.

ACKNOWLEDGEMENTS

To de Mexican Council of Science and Technology (CONACYT) for providing scholarship to Guerrero-Rangel for carrying out this research as a part of her Master's thesis.

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